

## basicity

For Brønsted bases, the tendency of a compound to act as hydron (proton) acceptor. The basicity of a chemical species is normally expressed by the acidity of the conjugate acid (see conjugate acid–base pair). For Lewis bases it relates to the association constants of Lewis adducts and  $\pi$ -adducts.

**Source:**

PAC, 1994, 66, 1077 (*Glossary of terms used in physical organic chemistry (IUPAC Recommendations 1994)*) on page 1088